



GOLDEN DAYS UNIVERSAL SCHOOL

ASSIGNMENT-1 (2020-21)

Class -X Div _____

SUBJECT- SCIENCE

Date / /2020

Q) SHORT ANSWER TYPE QUESTIONS:

1) Identify the following changes as chemical or physical:

- a) crumpling a sheet of aluminum foil b) Baking a cake c) Shredding paper
d) sublimation of dry ice e) changing colour f) Burning of paper

2) Write the chemical equation of the reaction in which the following changes have taken place with an example:

- a) Change in colour b) Change in temperature c) formation of precipitation

3) What is meant by skeletal type chemical equation? What does it represent? Using the equation for electrolytic decomposition of water, differentiate between a skeletal chemical equation and a balanced chemical equation.

4) Write chemical equation for the reactions taking place when,

- a) iron reacts with steam b) Magnesium reacts with dil. HCl c) copper is heated in air

5) Write balanced equation for the following mentioning the type of chemical reaction involved:

- a) Aluminium + Bromine $\rightarrow$ Aluminium bromide
b) Calcium carbonate $\rightarrow$ Calcium oxide + Carbon dioxide
c) Silver chloride $\rightarrow$ Silver + Chlorine

6) Write balanced chemical equation for the reaction that takes place during respiration. Identify the type of combination reaction that takes place during this process and justify the name. Give one more example of this type of reaction.

7) 2g of ferrous sulphate crystals are heated in a dry boiling tube.

- a) list any two observations b) name the type of reaction c) write the chemical equation for the reaction

8) Decomposition reaction require energy either in the form of heat or light or electricity for breaking down the reactants. write one equation each for decomposition reaction where energy is in the form of heat, light, electricity.

9) Balance the following chemical equations:

- 1) $\text{BaCl}_2 + \text{H}_2\text{SO}_4 \rightarrow \text{BaSO}_4 + \text{HCl}$ 2) $\text{Ca}(\text{OH})_2 + \text{HNO}_3 \rightarrow \text{Ca}(\text{NO}_3)_2 + \text{H}_2\text{O}$
3) $\text{Pb}(\text{NO}_3)_2 \rightarrow \text{PbO} + \text{NO}_2 + \text{O}_2$ 4) $\text{MnO}_2 + \text{HCl} \rightarrow \text{MnCl}_2 + \text{H}_2\text{O} + \text{Cl}_2$

10) Write the balanced equation for the given reaction and identify the type of reaction :

Potassium bromide + Barium iodide $\rightarrow$ Potassium iodide + Barium bromide